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# Chemistry Chapter 9 Stoichiometry

**chapter 9 chemical calculations and chemical formulas** - chapter 9 - chemical calculations and chemical formulas 119 chapter 9 map chapter checklist read the review skills section. if there is any skill mentioned that you have not yet mastered, review the material on that topic before reading this chapter. read the chapter quickly before the lecture that describes it. **chapter 9 - covalent bonding: orbitals - sciencegeek** - chapter 9 - covalent bonding: orbitals . 9.1 hybridization and the localized electron model . a. hybridization 1. the mixing of two or more atomic orbitals of similar energies on the same atom to produce new orbitals of equal energies b. hybrid orbitals 1. orbitals of equal energy produced by the combination of two or more orbitals on the same atom **chapter 9 practice test - naming and writing chemical formulas** - chapter 9 practice test - naming and writing chemical formulas matching match each item with the correct statement below. match each item with the correct statement below. a. monatomic ion f. cation b. acid g. binary compound c. base h. anion d. law of definite proportions i. polyatomic ion e. law of multiple proportions \_\_\_\_ 1. **chapter 9 chemical c and c formulas - mark bishop** - chapter 9 chemical calculations and chemical formulas 329 n order to explore and make use of the seemingly limitless changes that matter can undergo, chemists and chemistry students often need to answer questions that **chemistry chapter 9 outline notes** - chemistry chapter 9 outline notes covalent bonding 9.1 - the covalent bond o why do atoms bond? o what is a covalent bond? o single covalent bonds • sigma bond o multiple covalent bonds • pi bonds o strength of covalent bonds 9.2 - naming molecules (covalent compounds) o naming binary molecular compounds • prefixes o naming acids **chapter 9 practice test - welcome to mrs. fischer's site!** - chapter 9 practice test multiple choice identify the choice that best completes the statement or answers the question. \*\*\*\*you need to know all the rules for naming ionic, molecular and acidic compounds. \*\*\*\*you will have questions over ch 1,2,4 and 25 as well. **chapter 9 chemical quantities - francis howell high school** - e. molar mass of  $\text{Fe}(\text{NO}_3)_3 = 241.9 \text{ g}$  272 g  $\text{Fe}(\text{NO}_3)_3$  1 mol 241.9 g = 1.12 mol  $\text{Fe}(\text{NO}_3)_3$  f. molar mass of  $\text{FeSO}_4 = 151.9 \text{ g}$  12.7 mg  $\text{FeSO}_4$  1 g 1000 mg 1 mol 151.9 g = 8.36  $10^{-5}$  mol  $\text{FeSO}_4$  g. molar mass of  $\text{LiOH} = 23.95 \text{ g}$  6.91 103 g  $\text{LiOH}$  1 mol 23.95 g = 288.5 = 289 mol  $\text{LiOH}$  14. a. molar mass of  $\text{CaCO}_3 = 100.1 \text{ g}$  2.21  $10^{-4}$  mol  $\text{CaCO}_3$  100.1 g 1 mol = 0 ... **name date class chemical names and formulas 9** - chapter 9, chemical names and formulas(continued) 8. use the periodic table to write the name and formula (including charge) for each ion in the table below. polyatomic ions (pages 257-258) 9. what is a polyatomic ion? 10. is the following sentence true or false? the names of polyatomic anions always end in -ide. \_\_\_\_ 11.

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